

3-(4-Fluorophenyl)-1,5-diphenyl-2-pyrazoline

Huan-Mei Guo,^a Fang-Fang
Jian,^{b*} Li-You Zhou,^c Pu-Su
Zhao^b and Jian Zheng^b^aDepartment of Chemistry, Weifang College,
Weifang 261061, People's Republic of China,^bNew Materials and Function Coordination
Chemistry Laboratory, Qingdao University of
Science and Technology, Qingdao 266042,
People's Republic of China, and ^cDepartment of
Physics, Weifang College, Weifang 261061,
People's Republic of China

Correspondence e-mail: zhaopu@163.com

Key indicators

Single-crystal X-ray study

T = 293 K

Mean $\sigma(\text{C}-\text{C}) = 0.003 \text{ \AA}$

R factor = 0.050

wR factor = 0.156

Data-to-parameter ratio = 16.5

For details of how these key indicators were
automatically derived from the article, see
<http://journals.iucr.org/e>.

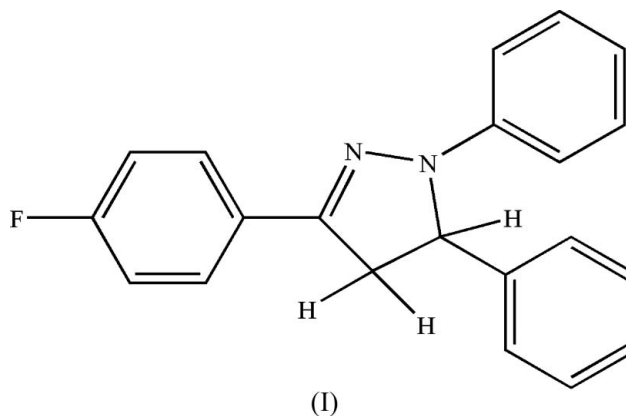
The title compound, $\text{C}_{21}\text{H}_{17}\text{FN}_2$, crystallizes with two independent molecules in the asymmetric unit. The pyrazoline and 4-fluorophenyl rings are planar. The phenyl ring at the 5-position and the pyrazoline heterocycle are almost perpendicular.

Received 9 August 2006

Accepted 4 September 2006

Comment

Five- and six-membered heterocyclic compounds are important constituents that often exist in biologically active natural products and synthetic compounds of medicinal interest (Gilchrist, 1998). Among them, 1,3,5-trisubstituted pyrazolines can be easily prepared from phenylhydrazine and chalcone derivatives (Nakamichi *et al.*, 2002). Some 1,3,5-triaryl-substituted pyrazolines can also be used as probes in a biological environment (Fahrni *et al.*, 2003). Here, we report the crystal structure of such a compound, (I).



The structure of (I) (Fig. 1) contains two crystallographically independent molecules in the asymmetric unit, hereafter named *F1* and *F2*. In both molecules, the $\text{C}=\text{N}$ bond lengths are slightly longer than those found in similar structures [1.291 (2) \AA (Rurack *et al.*, 2000), 1.283 (2) \AA for polymorph α (Kimura *et al.*, 1977) and 1.293 (3)/1.291 (3) \AA (Ge, 2006)]. In contrast, the $\text{N1}-\text{N2}$ and $\text{N3}-\text{N4}$ bond lengths are shorter than those found in the above-cited structures [1.394 (3), 1.390 (3) and 1.384 (2)/1.385 (2) \AA , respectively]. The $\text{C}-\text{F}$ bond lengths are slightly shorter than those found in similar structures [*e.g.* 1.370 (2) \AA ; Joshi *et al.*, 1992].

All the bond lengths and angles for the phenyl and benzene rings in (I) are in normal ranges. In molecule *F1*, the pyrazoline ring and 4-fluorophenyl group define a plane (*P1*), the largest deviation being 0.068 (3) \AA for atom N2 . The dihedral angles between *P1* and the phenyl rings at positions 1 and 5 of the pyrazoline are 12.64 (2) and 78.95 (3)°, respectively. The

dihedral angle between the phenyl rings is 88.51 (2)°. Similarly, in molecule *F2*, the pyrazoline and 4-fluorophenyl rings also define a plane (*P2*), the largest deviation being 0.072 (3) Å for atom C42. The dihedral angles between *P2* and phenyl groups at positions 1 and 5 of the pyrazoline are 17.59 (2) and 75.29 (3)°, respectively. The dihedral angle between the two phenyl rings is 88.32 (3)°.

Experimental

The title compound was prepared by reaction of phenylhydrazine (0.02 mol) and 1-(4-fluorophenyl)-3-phenyl-2-propenyl-1-ketone (0.02 mol) dissolved in acetic acid (40 ml). Single crystals suitable for X-ray measurements were obtained by recrystallization from EtOH at 298 K.

Crystal data

$C_{21}H_{17}FN_2$	$V = 1678.2 (6) \text{ \AA}^3$
$M_r = 316.37$	$Z = 4$
Triclinic, $P\bar{1}$	$D_x = 1.252 \text{ Mg m}^{-3}$
$a = 11.491 (2) \text{ \AA}$	Mo $K\alpha$ radiation
$b = 11.728 (2) \text{ \AA}$	$\mu = 0.08 \text{ mm}^{-1}$
$c = 13.601 (3) \text{ \AA}$	$T = 293 (2) \text{ K}$
$\alpha = 89.44 (3)^\circ$	Block, colourless
$\beta = 71.07 (3)^\circ$	$0.23 \times 0.22 \times 0.19 \text{ mm}$
$\gamma = 76.07 (3)^\circ$	

Data collection

Enraf–Nonius CAD-4 diffractometer	3538 reflections with $I > 2\sigma(I)$
ω scans	$R_{\text{int}} = 0.030$
Absorption correction: none	$\theta_{\text{max}} = 27.2^\circ$
7526 measured reflections	3 standard reflections every 100 reflections
7160 independent reflections	intensity decay: <1%

Refinement

Refinement on F^2	$w = 1/[\sigma^2(F_o^2) + (0.0765P)^2 + 0.0731P]$
$R[F^2 > 2\sigma(F^2)] = 0.050$	where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2) = 0.156$	$(\Delta/\sigma)_{\text{max}} < 0.001$
$S = 1.01$	$\Delta\rho_{\text{max}} = 0.18 \text{ e \AA}^{-3}$
7160 reflections	$\Delta\rho_{\text{min}} = -0.16 \text{ e \AA}^{-3}$
434 parameters	Extinction correction: <i>SHELXL97</i>
H-atom parameters constrained	Extinction coefficient: 0.023 (2)

Table 1

Selected geometric parameters (Å, °).

F1–C19	1.366 (2)	F2–C40	1.357 (2)
N1–N2	1.374 (2)	N3–N4	1.378 (2)
N1–C6	1.396 (3)	N3–C27	1.396 (2)
N1–C7	1.469 (3)	N3–C28	1.477 (3)
N2–C15	1.298 (2)	N4–C36	1.295 (2)
C1–C2	1.379 (3)	C22–C23	1.376 (3)
C3–C4	1.368 (4)	C25–C26	1.386 (3)
C7–C8	1.531 (3)	C29–C30	1.374 (3)
C8–C9	1.368 (3)	C37–C38	1.399 (3)
C16–C21	1.372 (3)		
N1–C7–C8	112.05 (17)	N3–C28–C29	112.28 (17)

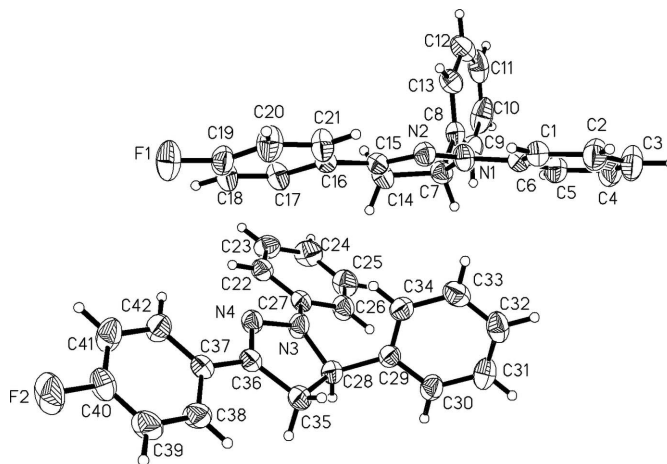


Figure 1

The asymmetric unit of (I), showing the atom-labelling scheme, with displacement ellipsoids drawn at the 30% probability level.

All H atoms were positioned geometrically and refined using a riding model, with C–H = 0.93 (aromatic CH), 0.97 (methylene CH₂) or 0.98 Å (methine CH). Isotropic displacement parameters for H atoms were fixed at $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

Data collection: *CAD-4 Software* (Enraf–Nonius, 1989); cell refinement: *CAD-4 Software*; data reduction: *NRCVAX* (Gabe *et al.*, 1989); program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997a); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997a); molecular graphics: *SHELXTL/PC* (Sheldrick, 1997b); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

The authors thank the Natural Science Foundation of Shandong Province (grant No. Y2005B04) and the Doctoral Fund of Qingdao University of Science and Technology.

References

- Enraf–Nonius (1989). *CAD-4 Software*. Version 5.0. Enraf–Nonius, Delft, The Netherlands.
- Fahrni, C. J., Yang, L. & VanDerveer, D. G. (2003). *J. Am. Chem. Soc.* **125**, 3799–3812.
- Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
- Gabe, E. J., Le Page, Y., Charland, J.-P., Lee, F. L. & White, P. S. (1989). *J. Appl. Cryst.* **22**, 384–387.
- Ge, W.-Z. (2006). *Acta Cryst.* **E62**, o3109–o3110.
- Gilchrist, T. L. (1998). *Heterocyclic Chemistry*, 3rd ed. London: Addison-Wesley Longman Ltd.
- Joshi, K. C., Bohra, R. & Joshi, B. S. (1992). *Inorg. Chem.* **31**, 598–603.
- Kimura, T., Kai, Y., Yasuoka, N. & Kasai, N. (1977). *Acta Cryst.* **B33**, 1786–1792.
- Nakamichi, N., Kawashita, Y. & Hayashi, M. (2002). *Org. Lett.* **4**, 3955–3957.
- Rurack, K., Bricks, J. L., Schulz, B., Maus, M., Reck, G. & Resch-Genger, U. (2000). *J. Phys. Chem. A*, **104**, 6171–6188.
- Sheldrick, G. M. (1997). *SHELXS97* and *SHELXL97*. University of Göttingen, Germany.
- Sheldrick, G. M. (1997). *SHELXTL/PC*. Bruker AXS Inc., Madison, Wisconsin, USA.